Chemistry 115 Name KEY

Dr. Cary Willard

Quiz 8A (20 points) April 13, 2009

1. (3 points) Draw a lewis electron dot structure for phosphorous?



1. (4 points) Draw a lewis electron dot structure for carbon tetrabromide, CBr4.



1. (4 points) Draw a lewis electron dot structure for ethyne, C2H2.



1. (3 points) Why are only valence electrons represented in a lewis structure?

Because only the outermost valence electrons are involved in bonding.

1. (3 points) How do ionization energies change as you move across the periodic table to the right?

Ionization energies increase as you move across the periodic table to the right.

1. (3 points) Explain the reason for the trend you described in question 5.

As you move across the periodic table to the right, the effective nuclear increases due to an increase in the number of protons in the nucleus without much shielding of the charge by the electrons in that outer level.

Chemistry 115 Name KEY

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Quiz 8B (20 points) April 13, 2009

1. (3 points) Draw a lewis electron dot structure for sulfur?



1. (4 points) Draw a lewis electron dot structure for carbon tetrachloride, CCl4.



1. (4 points) Draw a lewis electron dot structure for ethyne, C2H2.



1. (3 points) Why are only valence electrons represented in a lewis structure?

Because only the outermost valence electrons are involved in bonding.

1. (3 points) How do ionization energies change as you move across the periodic table to the right?

Ionization energies increase as you move across the periodic table to the right.

1. (3 points) Explain the reason for the trend you described in question 5.

As you move across the periodic table to the right, the effective nuclear increases due to an increase in the number of protons in the nucleus without much shielding of the charge by the electrons in that outer level.